

# The Physical Chemistry Of Natural Waters

Frank J Millero

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The physical chemistry of water has advanced considerably because of the instrumentation that is available to study it. Properties of water are based in large part on how the individual molecules interact with their nearest neighbors. The nature of liquid water is determined by the hydrogen bond and van der Waals forces (Schmid, 2001). The definition used by advanced physical chemists is more complex than what is taught in introductory chemistry classes. These properties may seem esoteric to aquatic ecology students, but they do determine how processes like weathering and transport of nonpolar pollutants occur. Read full chapter. The natural environment also differs because of the large number of chemical species that coexist and interact concurrently.